Study Guide for Unit 2 lesson 4-5 quiz

Name:\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_Date:\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_Period:\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_

1. How are atoms arranged on a periodic table?

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|  | Metalloids | metals | Non-metals |
| Properties of elements |  |  |  |

1. What three particles make up atoms?
2. Which particle has a negative charge?
3. Which particles are found in the nucleus of an atom?
4. Differentiate a pure substance from a mixture.
5. How many neutrons are found in atom whose atomic number is 8 and mass is 16?
6. What would the atomic number be for an atom which has 5 protons, 5 electrons, and 6 neutrons?
7. How are isotopes of an atom different from one another?
8. What are the rows in a periodic table called? What are the columns called?
9. How many valence electrons does element P have? How many of those valence electrons can be bonded to a different atom?
10. How are compounds different from elements?
11. Create a bohr model of Ar. Label the valence electrons.
12. If atoms share the same number of protons what could you infer about those two atoms?
13. How many neutrons does an atom whose atomic number is 22 have?
14. How many electrons can be found in the nucleus of an atom with an atomic number of 15?
15. List in order from the least spread out to the most spread out atoms for the three states of matter below

Liquid, solid, gas

1. Explain ductility and what type of property it is.